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Answer. Molar mass = 151.1 g/mol Potassium: $(39.10 / 101.1) \times 100 = 38.67\%$ Nitrogen: $(14.01 / 101.1) \times 100 = 13.86\%$ Oxygen: $(48.00 / 101.1) \times 100 = 47.48\%$. Molar mass = 164.1 g/mol Potassium: $(39.10 / 101.1) \times 100 = 38.67\%$ Nitrogen: $(14.01 / 101.1) \times 100 = 13.86\%$ Oxygen: $(48.00 / 101.1) \times 100 = 47.48\%$.

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FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas
they contain. Hydrogen molecule 1 mol H₂ at STP = 22.4 L Oxygen molecule 1 mol O₂ at STP = 22.4 L Carbon dioxide molecule 1 mol CO₂
at STP = 22.4 L CHAPTER 11 Molecular Composition of Gases Start

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Atoms, Molecules, and Ions Jim Geiger Cem 151 Atoms, Molecules, and Ions Law of Constant Composition Joseph Proust (1754–1826)

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